Titration Reinforcement (S332.7.7) Name \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Chem 332 – O’Dette Date \_\_\_\_\_\_\_\_\_\_\_\_ Period \_\_\_\_

1. A student titrates 40.00 mL of a HCl solution of unknown concentration with a 0.5500 M NaOH solution. The volume of base solution needed to reach the equivalence point is 24.64 mL. What is the molarity of the HCl solution? (0.3388 M)
2. If 50.00 mL of a 1.000 M H2SO4 solution neutralizes 35.4 mL of KOH, what is the molarity of the KOH solution? (2.82 M)
3. What is the molarity of sodium hydroxide if 20.0 mL of the solution is neutralized by 17.4 mL of 1 M H3PO4? (2.61 M)
4. What is the molarity of 25.7mL of Ca(OH)2 if the solution requires 15.0 mL of 2.0 M H3PO4? (1.8 M)