pH Calculations Reinforcement (S332.7.6) Name \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Chem 332 – O’Dette Date \_\_\_\_\_\_\_\_\_\_\_\_ Period \_\_\_\_

1. Calculate the [H+] for a solution with a pH of 12.8.
2. Calculate the pH of a solution with a [OH-] = 4.3x10-5 M.
3. Calculate the pH of an aqueous solution that contains 5.00 g of HCl in 200 mL of solution.
4. What is the molarity of a solution of the base magnesium hydroxide that has a pH of 9.7?
5. The pH of a citric acid solution is 3.15. What are the [H+] and [OH-] in this solution?

Answers: 1) 1.6x10-13 M 2) 9.6 3) 4) 5) [H+] = 7.1x10-4 M [OH-] = 1.4x10-11 M