Acid Base Theories Reinforcement (S332.7.4) Name \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Chem 332 – O’Dette Date \_\_\_\_\_\_\_\_\_\_\_\_ Period \_\_\_\_

1. How did Arrhenius define an acid and a base?
2. How are acids and bases defined by the Bronsted-Lowry theory?
3. Classify each compound as an Arrhenius acid or an Arrhenius base.
	1. Ca(OH)2
	2. HNO3
	3. KOH
	4. C2H5COOH
	5. HBr
	6. H2SO4
4. Use the Bronsted-Lowry definitions of acids and bases to identify each reactant as an acid or a base.
	1. KOH + HBr KBr + H2O
	2. HCl + H2O Cl- + H3O+