Acid Base Theories Reinforcement (S332.7.4) Name \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Chem 332 – O’Dette Date \_\_\_\_\_\_\_\_\_\_\_\_ Period \_\_\_\_

1. How did Arrhenius define an acid and a base?
2. How are acids and bases defined by the Bronsted-Lowry theory?
3. Classify each compound as an Arrhenius acid or an Arrhenius base.
   1. Ca(OH)2
   2. HNO3
   3. KOH
   4. C2H5COOH
   5. HBr
   6. H2SO4
4. Use the Bronsted-Lowry definitions of acids and bases to identify each reactant as an acid or a base.
   1. KOH + HBr KBr + H2O
   2. HCl + H2O Cl- + H3O+