Molarity Reinforcement (11.1) Name \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Chem 332 – O’Dette Date \_\_\_\_\_\_\_\_\_\_\_\_ Period \_\_\_\_

1. Vinegar contains 5.0 g of acetic acid, CH3COOH, in 100.0 mL of solution. What is the molarity of the vinegar?
2. What mass of AgNO3 is needed to prepare 250.0 mL of a 0.125 M solution?
3. You have 1 L of 1 M NaCl, and 1 L of 1 M KCl. Which solution has the greater mass of solute?
4. One solution contains 55 g NaCl per liter, and another contains 55 g KCl per liter. Which solution has the higher molarity? How can you tell?
5. How many milliliters of 1.0 M AgNO3 are needed to provide 168.88 g of AgNO3?

Answers: 1) 0.83 M 2) 5.30 g 3) KCl 4) NaCL; there are mole moles of NaCl in 55 g 5) 994 mL