Percent Yield Reinforcement (S332.5.6) Name \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Chem 332 – O’Dette Date \_\_\_\_\_\_\_\_\_\_\_\_ Period \_\_\_\_\_\_

1. What is the equation used to find percent yield? Define the variables.
2. When 84.8 g of iron(III) oxide reacts with carbon monoxide, iron is produced.

Fe2O3 + 3 CO 2 Fe + 3 CO2

What is the theoretical yield of iron? What is the percent yield if 52.6 g of Fe are actually produced?

1. When 5.00 g of copper reacts with silver nitrate, silver metal and copper(II) nitrate are produced. What is the theoretical yield of silver in this reaction?
2. What is the percent yield if 4.65 g of copper is produced when 1.87 g of aluminum reacts?

2 Al + 3 CuSO4 Al2(SO4)3 + 3 Cu

Answers: 2)59.3 g; 88.7% 3) 17.0 g 4)70.5%