Covalent Bond Reinforcement (S332.3.2) Name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Chem 332 – O’Dette Period: \_\_\_\_ Date: \_\_\_\_\_\_\_\_\_\_

1. What are the properties of a covalent bond?
2. Do covalent bonds share or transfer electrons?
3. Are all covalent compounds soluble in water? Explain
4. Place a check mark next to those elements that would form a covalent bond.
   1. H + Cl \_\_\_\_\_\_\_
   2. S + Br \_\_\_\_\_\_\_\_\_\_
   3. Ca + Cl \_\_\_\_\_\_\_\_\_
   4. Te + F \_\_\_\_\_\_\_\_\_
   5. N + Br \_\_\_\_\_\_\_\_\_\_
   6. K + F \_\_\_\_\_\_\_\_\_
   7. P + F \_\_\_\_\_\_\_\_\_\_
   8. C + Rn \_\_\_\_\_\_\_\_
5. True or False. Molecular compounds have relatively high boiling and melting points.
6. True or False. Molecular compounds cannot conduct electricity in the solid state.
7. True or False. Molecular compounds can be soluble in both water and hexane.