Ionic Bond Reinforcement (S332.3.1) Name \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Chem 332 – O’Dette Date \_\_\_\_\_\_\_\_\_\_\_\_\_\_ Period \_\_\_\_\_

Directions: Please answer the following questions.

1. How can you tell if a compound is ionic?
2. Explain why ionic compounds are electrically neutral.
3. Which of the following pairs of atoms would you expect to combine chemically to form an ionic compound?
   1. Li and S
   2. O and S
   3. Al and O
   4. F and Cl
   5. I and K
   6. H and N
4. Which of the following compounds have an ionic bond?

NaCl, HCl, H2O, KOH, F2

1. Do ionic bonds share or transfer electrons?
2. What are properties of an ionic bond?

For the following questions, mark them as True(T) or False(F).

1. Ionic compounds cannot conduct electricity if they are dissolved in water. \_\_\_\_\_\_\_\_\_
2. Ionic compounds have relatively high boiling and melting points. \_\_\_\_\_\_\_
3. Ionic compounds do not conduct electricity in the solid state. \_\_\_\_\_\_\_\_\_