Ionization Energy Reinforcement (S332.2.12) Name \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Chem 332 – O’Dette Period \_\_\_\_\_\_

1. Define ionization energy in your own words. What is it affected by?
2. Circle the element in each pair that has the highest ionization energy:

a) Na or Cl b) Na or Cs c) F or I d) K or F e) Mg or S f) N or Sb

1. Circle the element in each pair that has the lowest ionization:

a) Na or Cl b) Na or Cs c) F or I d) K or F e) Mg or S f) N or Sb

1. When elements have a large radius, they tend to have a (large, small) ionization energy. WHY???
2. Which of the alkaline earth metals has the smallest ionization energy?