Atom Trends Reinforcement (S332.2.11) Name \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Chem 332 – O’Dette Period \_\_\_\_\_\_\_

1. Place the following elements in order of decreasing atomic mass: Ar, K, Cl, S, Ca.
2. Which has the largest radius: O or F? Briefly explain your answer.
3. What affects the atomic radius? In other words, how is it possible for it to change within the same element?
4. Circle the element in each pair that has the largest atomic radius:

a) K or Br b) F or Br c) He or Rn d) Mg or Cl e) O or S f) Be or O

Now circle the element in each pair that has the largest atomic mass:

a) K or Br b) F or Br c) He or Rn d) Mg or Cl e) O or S f) Be or O

1. Which element in the second period has the greatest atomic mass?
2. Which of the metalloids has the smallest atomic radius?