Acid and Base Notes Name \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Chem 332 – O’Dette Date \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Period \_\_\_\_

Water

* When a water molecule \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_, it forms two ions.
* Draw the water molecule splitting

Acid Base Theory

1. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
	* Svante Arrhenius in 1884 established the first theory of acids and bases
	* Acids produce \_\_\_\_\_\_ when dissolved in water
	* Bases produce \_\_\_\_\_\_ when dissolved in water
	* EX:
2. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
	* Arrhenius theory doesn’t always work, for example NH3 would be considered an acid by Arrhenius theory but when NH3 reacts it will actually act as a base
	* In 1923 Johannes Bronsted and Thomas Lowry proposed a new acid base theory
	* Acid – \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
	* Base – \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
	* EX:

Strength of Acids and Bases

* Depends on \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
* Dissociation =
	+ Acid Ex:
	+ Base Ex
* Strong Acids :
* Weak Acids :
* Strong Bases :
* Weak Bases :

Properties of Acids and Bases

* Acids :
* Bases: